

Unit 0 - Matter and Measurement

Term	Definition
Accuracy	How close your results are to the desired value
Chromatography	Separation of particles based on size and solubility.
Compound	Two or more elements chemically combined. Formula which contains 2+ different symbols/elements
Density	Amount of mass in a given space; RATIO of mass to volume
Distillation	Separate 2 or more liquids based on their boiling points.
Element	Simplest form of matter. Found on the periodic table.
Extensive Property	A property that depends on how much material you are dealing with. Can not be used to help identify matter
Filtration	Separation of substances based on solubility
Gram	Basic SI unit of mass
Heterogeneous Mixture	A mixture where the substances aren't equally distributed. Different throughout.
Homogeneous Mixture	A mixture that is the same throughout.
Intensive	A property that <i>does not</i> depend on how much material you are dealing with. Can be used to help identify matter; a constant about that particular type of matter.
Liter	Basic SI unit of Volume
Mass	The amount of matter in an object. The more mass, the more stuff is present.
Matter	Anything that has mass and volume (takes up space)
Meter	Basic SI unit of distance/length
Mixture	2 or more substances combined physically. Not a pure substance.
Particle Diagram	A drawing that represents atoms or molecules
Precision	A measurement of how repeatable a measurement is. The more significant figures, the more precise the measurement.
Pure Substance	A substance where each particle has the same composition
S.I. unit	The modern form of the metric system and is the world's most widely used system of measurement
Scientific Notation	Method for expressing very large or small numbers easily (Example: 6.02×10^{23} atoms = 1 mole)
Significant Figures	The number of digits in a number that tell you useful information.
Volume	Amount of SPACE an object takes up
Weight	The effect of gravity on an object's mass.