

Term	Definition
Binary compound	compound containing 2 ELEMENTS (example: CO)
Bond	forces of attraction that hold atoms together in a molecule or compound
Compound	a substance composed of two or more atoms from different elements CHEMICALLY bonded together
Covalent Bond	chemical bond involving the SHARING of electrons between two nonmetal atoms; electronegativity difference between elements typically LESS than 1.7
Dipole	a difference in charge between 2 atoms in a bond
Dissociate	a process in which molecules separate or split into smaller particles such as ions, usually in a reversible manner.
Endothermic	energy is CONSUMED as a product of a chemical reaction
Exothermic	energy is RELEASED as a product of a chemical reaction
Intermolecular forces (IMF's)	weak forces between molecules that hold the molecules to one another; not actually chemical bonds
Ionic Bond	chemical bond involving the TRANSFER of electrons between a metal and nonmetal atom (metals lose, nonmetals gain); electronegativity difference between elements typically GREATER than 1.7
Molecule	a COVALENTLY bonded substance; can be atoms of the same element
Nonpolar molecules	a molecule with equal sharing of electrons; a symmetrical covalent molecule
Octet Rule	atoms bond together in order to have 8 electrons in their valence shell
Oxidation number	the "charge" an element has within a compound
Polar molecule	a covalent molecule with an unequal sharing of electrons; an asymmetrical covalent molecule
Polyatomic ions	atoms of two or more elements chemically bonded together and having a NET CHARGE
Stock system	system using Roman numerals (appearing in parentheses) after the element symbol to indicate the oxidation number of a TRANSITION METAL
Ternary compound	compound containing 3 ELEMENTS (example: C ₆ H ₁₂ O ₆)